

ASSIGNMENT MODULE C

1. List the fractional concentrations of the four major gases that comprise the atmosphere

Gas	Fractional Concentration
A. Nitrogen	78%
B. Oxygen	21%
C. Argon	0.9%
D. Carbon Dioxide	0.03%

2. Given a barometric pressure of 760 torr, calculate the P_{iO_2}
 $(P_{Baro} - 47 \text{ mm Hg}) \times F_{iO_2} = (760-47) \times .21 = 150 \text{ mm Hg}$
REMEMBER TO SUBTRACT OUT 47 mm Hg WHENEVER YOU SEE THE SMALL "i" IN THE QUESTION (P_{iO_2} , F_{iO_2})
3. Given a barometric pressure of 760 torr, calculate the P_{iN_2}
 $(P_{Baro} - 47 \text{ mm Hg}) \times F_{iN_2} = (760-47) \times .78 = 556 \text{ mm Hg}$
4. Given a barometric pressure of 760 torr, calculate the P_{iAr}
 $(P_{Baro} - 47 \text{ mm Hg}) \times F_{iAr} = (760-47) \times .009 = 6.4 \text{ mm Hg}$
5. Given a barometric pressure of 760 torr, calculate the P_{iCO_2}
 $(P_{Baro} - 47 \text{ mm Hg}) \times F_{iCO_2} = (760-47) \times .0003 = 0.21 \text{ mm Hg}$
6. Given a barometric pressure of 750 torr, calculate the P_{iO_2}
 $(P_{Baro} - 47 \text{ mm Hg}) \times F_{iO_2} = (750-47) \times .21 = 148 \text{ mm Hg}$
7. Given a barometric pressure of 740 torr, calculate the P_{iCO_2}
 $(P_{Baro} - 47 \text{ mm Hg}) \times F_{iCO_2} = (740-47) \times .21 = 0.21 \text{ mm Hg}$
8. Given a barometric pressure of 755 torr and an F_{iO_2} of .40, calculate the P_{iO_2}
 $(P_{Baro} - 47 \text{ mm Hg}) \times F_{iO_2} = (755-47) \times .40 = 283 \text{ mm Hg}$
9. Given a barometric pressure of 735 torr and an F_{iO_2} of .60, calculate the P_{iO_2}
 $(P_{Baro} - 47 \text{ mm Hg}) \times F_{iO_2} = (735-47) \times .60 = 413 \text{ mm Hg}$
10. Given a barometric pressure of 740 torr and an F_{iO_2} of .30, calculate the P_{iO_2}
 $(P_{Baro} - 47 \text{ mm Hg}) \times F_{iO_2} = (740-47) \times .30 = 208 \text{ mm Hg}$
11. Given the following information, calculate the partial pressure of Gas C

Total Pressure in the gas mixture:	740 mm Hg
Partial Pressure of Gas A	400 torr
Partial Pressure of Gas B	30 torr
Partial Pressure of Gas D	50 torr
Partial Pressure of Gas E	35 torr
Partial Pressure of Gas C	?

$$P_{TOTAL} = P_A + P_B + P_C + P_D + P_E$$

$$P_C = P_{TOTAL} - (P_A + P_B + P_D + P_E) = 740 - 400 - 30 - 50 - 35 = 225 \text{ mm Hg}$$

12. Given a PB of 760 mm Hg, FiO_2 of .40 and a PaCO_2 of 50 torr, calculate the PAO_2 .

$$\text{PAO}_2 = [(\text{PBaro} - 47) \times \text{FiO}_2] - (\text{PaCO}_2 \times 1.25) = [(760-47) \times .4] - (50 \times 1.25) = [(713) \times (.4)] - 62.5 = 285.2 - 62.5 = 223 \text{ mm Hg}$$

13. Given a PB of 750 mm Hg, FiO_2 of .50 and a PaCO_2 of 70 torr, calculate the PAO_2 .

$$\text{PAO}_2 = [(\text{PBaro} - 47) \times \text{FiO}_2] - (\text{PaCO}_2 \times 1.25) = [(750-47) \times .5] - (70 \times 1.25) = [(703) \times (.5)] - 87.5 = 351.5 - 87.5 = 264 \text{ mm Hg}$$